

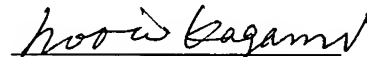


Japanese Patent Application No. 2004-110630

D E C L A R A T I O N

I, Norio KAGAMI, residing at Akasaka Office Heights, 13-5, Akasaka 4-chome, Minato-ku, Tokyo, Japan, do solemnly and sincerely declare that I have translated the Japanese Patent Office Priority Certificate and the certified copy of Japanese Patent Application No. 2004-110630, filed on April 5, 2004, based on which International Patent Application No. PCT/JP2005/005936 was filed on March 29, 2005, followed by U.S. Patent Application Serial No. 10/549,537 and attach a copy of my translation herewith and certify that the following is a true and faithful translation to the best my knowledge and belief.

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further, that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code, and that such willful false statements may jeopardize the validity of the application or any patent issuing thereon.



Norio KAGAMI

Patent Attorney

Dated this 11th day of June, 2007

JAPAN PATENT OFFICE

This is to certify that the annexed is a true copy of
the following application as filed with this Office.

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Application Number: Japanese Patent Application No.
2004-110630

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Applicant(s): NIKKO MATERIALS CO., LTD.

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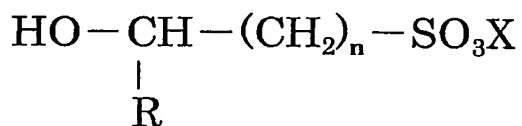
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[Designation of the Document] Claims

[Claim 1] An electroless gold plating liquid comprising a non-cyanide gold salt as a gold salt, an alkali metal salt or an ammonium salt of sulfurous acid and thiosulfuric acid as a complexing agent for gold, a hydroxyalkylsulfonic acid or a salt thereof represented by the following general formula as a reducing agent, and an amine compound.

[Chemical Formula 1]



(wherein R represents hydrogen, a carboxyl group, or any of a phenyl group, a naphthyl group, a saturated or unsaturated alkyl group, an acetyl group, an acetonyl group, a pyridyl group and a furyl group which may have a substitutional group, X represents any of hydrogen, Na, K, and NH₄, and n is an integer between 0 and 4.)

[Claim 2] An electroless gold plating liquid according to Claim 1, wherein the reducing agent represented by said general formula is sodium hydroxymethanesulfonate (HOCH₂SO₃Na).

[Claim 3] An electroless gold plating liquid according to Claim 1 or 2, wherein the non-cyanide gold salt is sodium gold sulfite.

[Claim 4] An electroless gold plating liquid according to any one of Claims 1 through 3, wherein said amine compound is ethylenediamine or glycine.

[Claim 5] A gold plated product produced by using an electroless gold plating liquid according to any one of Claims 1 through 4.

[Designation of the Document] Specification

[Title of the Invention] ELECTROLESS GOLD PLATING LIQUID

[Technical Field]

[0001]

The present invention relates to plating technology and specifically to an electroless gold plating liquid to obtain a gold plated film generally with an Au thickness of 0.4 μm or more, which is required mainly in gold wire bonding and TAB.

[Background Art]

[0002]

As electronic parts, electronic devices and the like trend to higher densities, electroless gold plating is coming into broader use, particularly since it is applicable to fine wiring. Reasons for this include the fact that circuits required for electrolytic plating around the object to be plated aren't needed because it is electroless, and the fact that it contributes to low-cost, simplified processes.

[0003]

However, cyanide baths have been used conventionally for electroless gold plating liquids because of their stability, raising concerns about damages on the resist. Moreover, the high toxicity of cyanide electroless gold plating liquids has led to demands for non-cyanide processes, but baths such as sulfite baths that do not contain cyanide, for example, tend to decompose easily, and improvements are needed. Under these circumstances, non-cyanide electroless gold plating liquids have already been proposed.

[0004]

Japanese Patent Documents 1 to 3 disclosed that an effective way for both bath stability and gold deposition speed of a non-cyanide electroless gold plating liquid is to use a mixed sulfite-thiosulfate complex system as a gold complexing agent and to use an urea compound such as thiourea

as a reducing agent. However, when thiourea is used as a reducing agent, it is oxidized and a non-water-soluble cyanamide polymer is produced, causing problems of plate deposition failure. Therefore, as effective countermeasures, addition of amine and carboxylic acid components for the aim of water-solubilizing the oxidation intermediate and addition of aliphatic saturated alcohol to suppress autolysis of the thiourea reducing agent are proposed.

[0005]

Moreover, Patent Document 4 states that addition of a mercaptobenzothiazole compound with the aim of suppressing contamination by metal impurities and improving bath stability is useful when using the aforementioned mixed complexing agent together with ascorbic acid as the reducing agent, however, reduction efficiency by ascorbic acid is poor, which needs to be added in excess to ensure a practical deposition speed (Patent Document 5). Patent Document 5 suggests the complexing agent system as described above, an aromatic compound typified by hydroquinone as reducing agent system, and heavy metal compound such as thallium nitrate in order to improve deposition speed, but, the use of heavy metals such as thallium is desired to be substituted with other safer substances from the view point of toxicity like cyanide.

[Patent Document 1] Japanese Patent Publication
No. H9-287077

[Patent Document 2] Japanese Patent Publication
No. H5-78854

[Patent Document 3] Japanese Patent Publication
No. H11-12753

[Patent Document 4] Japanese Patent No. 3148428

[Patent Document 5] Japanese Patent Publication
No. 2003-268559

[Disclosure of the Invention]

[Problems that the Invention Is to Solve]

[0006]

It is an object of the present invention to provide an electroless gold plating liquid which has an adequate deposition speed for practical use without containing thallium or other heavy metal ions, and which does not require a complicated bath composition or control such as addition of a water-solubilizing agent or autolysis-suppressing agent as thiourea reducing agents for example. It is also an object of the present invention to provide an electroless gold plating liquid which has an adequate deposition speed for practical use without containing thallium or other heavy metal ions, and which has better plating liquid stability than thiourea reducing agents.

[Means for Solving the Problems]

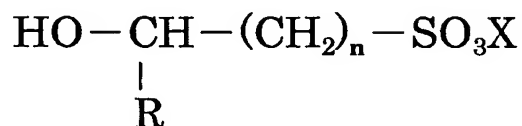
[0007]

After exhaustive research aimed at solving the aforementioned problems, the present invention was achieved when it was discovered that excellent stability of the plating liquid and adequate deposition speed for practical use could be achieved by including, in an electroless gold plating liquid using a non-cyanide gold salt, an alkali metal salt or an ammonium salt of sulfurous acid and thiosulfuric acid as a complexing agent for gold, a specific hydroxyalkylsulfonic acid or a salt thereof as a reducing agent, and an amine compound. Thereupon the present invention is as follows:

[0008]

(1) An electroless gold plating liquid comprising a non-cyanide gold salt as a gold salt, an alkali metal salt or an ammonium salt of sulfurous acid and thiosulfuric acid as a complexing agent for gold, a hydroxyalkylsulfonic acid or a salt thereof represented by the following general formula as a reducing agent, and an amine compound.

[Chemical Formula 1]



(wherein R represents hydrogen, a carboxyl group, or any of a phenyl group, a naphthyl group, a saturated or unsaturated alkyl group, an acetyl group, an acetyl group, a pyridyl group and a furyl group which may have a substitutional group, X represents any of hydrogen, Na, K, and NH_4 ; and n is an integer between 0 and 4.)

[0009]

(2) An electroless gold plating liquid according to (1) above wherein the reducing agent represented by said general formula is sodium hydroxymethanesulfonate ($\text{HOCH}_2\text{SO}_3\text{Na}$).

(3) An electroless gold plating liquid according to (1) or (2) above, wherein the non-cyanide gold salt is sodium gold sulfite.

(4) An electroless gold plating liquid according to any one of (1) through (4) above, wherein said amine compound is ethylenediamine or glycine.

(5) It is related to a gold plated product produced by using an electroless gold plating liquid according to any one of (1) through (4) above.

[Effect of the Invention]

[0010]

By using the electroless gold plating liquid of the present invention, the plating liquid is extremely stable and the deposition rate is adequate for practical use.

[Best Modes for Carrying out the Invention]

[0011]

The electroless gold plating liquid of the present invention is explained in detail below.

There are no particular limits on the non-cyanide gold salt

used as the gold source in the present invention as long as it is water-soluble: the most important feature is to involve an alkali metal salt or an ammonium salt of sulfurous acid and thiosulfuric acid as the complexing agent for gold, the aforementioned specific hydroxyalkylsulfonic acid or a salt thereof as the reducing agent, and an amine compound.

[0012]

As the non-cyanide water-soluble gold salts, chloroaurates, gold thiosulfates, gold thiomalates, and gold sulfites are exemplified. Of these, a gold sulfite is used by preference. The plating liquid contains the non-cyanide water-soluble gold salt preferably at a gold concentration of 0.1 to 10 g/L, more preferably 3 to 5 g/L. If the gold concentration is less than 0.1 g/L, the gold deposition rate will be much slower, while there is no advantage to be more than 10 g/L because the effects reach saturation.

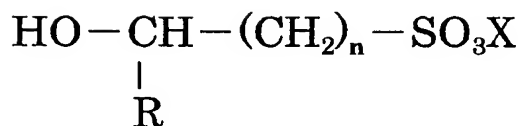
[0013]

A mixed complexing agent of sulfite-thiosulfate is used as the complexing agent for gold, and preferably 0.01 mol/L to 0.5 mol/L of the thiosulfuric acid salt is included, since at less than 0.01 mol/L the plating liquid is unstable and liable to decompose, while above 0.5 mol/L the effects on the plating reaction reach saturation. The content of the sulfurous acid salt is preferably 0.1 mol/L to 1.0 mol/L, since below 0.1 mol/L the plating bath is unstable and liable to decompose, while more than 1.0 mol/L is undesirable from a practical standpoint because the plating speed decreases.

[0014]

A hydroxyalkylsulfonic acid or a salt thereof represented by the following general formula is used as the reducing agent:

[Chemical Formula 2]



(wherein R represents hydrogen, a carboxyl group, or any of a phenyl group, a naphthyl group, a saturated or unsaturated alkyl group, an acetyl group, an acetonyl group, a pyridyl group and a furyl group which may have a substitutional group, X represents any of hydrogen, Na, K, and NH_4 , and n is an integer between 0 and 4.)

[0015]

In the aforementioned formula, examples of the substitutional group in the phenyl group, naphthyl group, saturated or unsaturated alkyl group, acetyl group, acetonyl group, pyridyl group and furyl group of R include a halogen atom, an alkoxy group, a nitro group, a hydroxyl group and a sulfonic acid group or a salt thereof and a phenyl group and an acetyl group and the like. Chlorine is desirable as the halogen atom, and a lower alkoxy group such as methoxy group is desirable as the alkoxy group. Salts of a sulfonic acid group include alkali metal salts or the like, and a sodium salt is preferred.

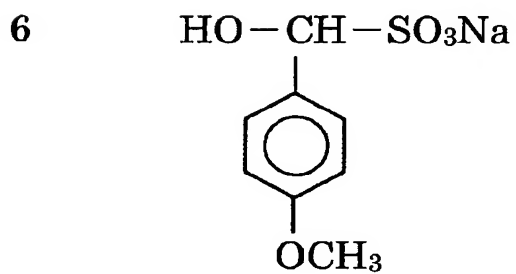
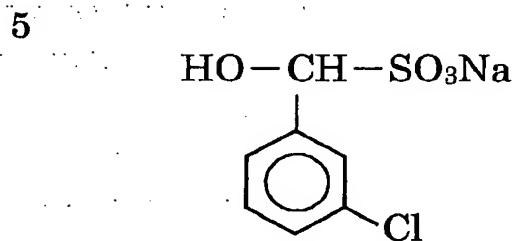
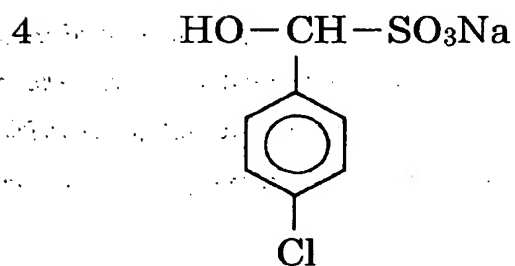
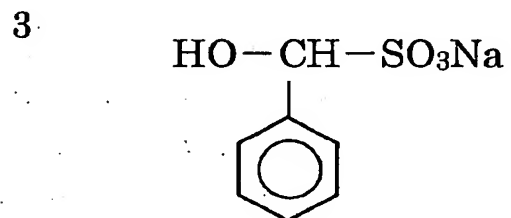
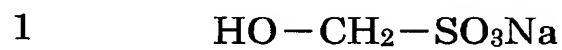
An alkyl group with 1 to 4 carbon atoms is desirable as the saturated or unsaturated alkyl group. Moreover, sodium is desirable as X in the above formula.

[0016]

Specific examples of the hydroxyalkylsulfonic acid or a salt thereof represented by the aforementioned general formula include the following compounds or the like for example, and in particular the compounds of No. 1 and No. 2 can be used by preference.

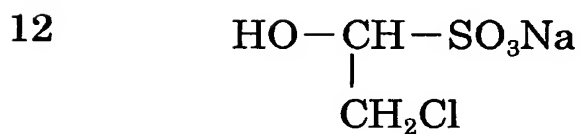
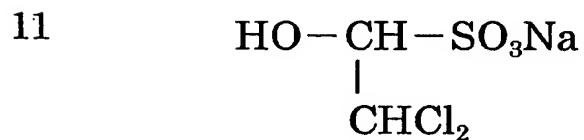
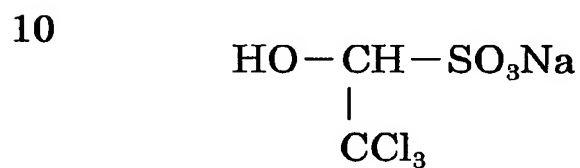
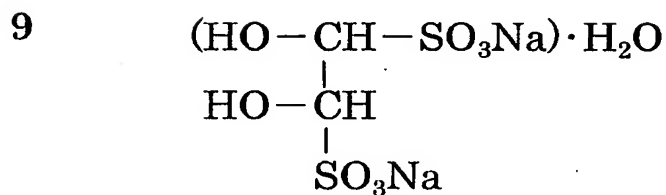
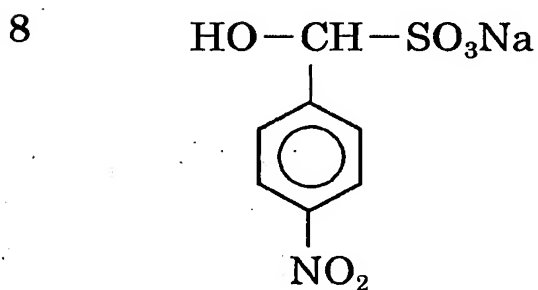
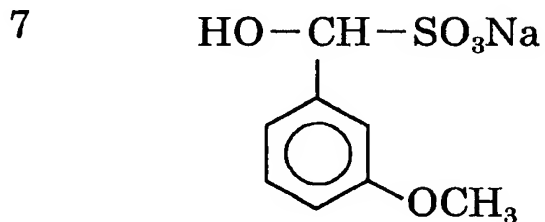
[0017]

[Chemical Formula 3]



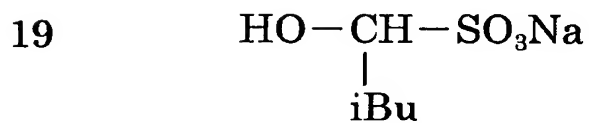
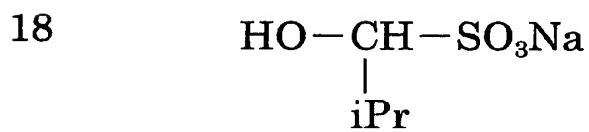
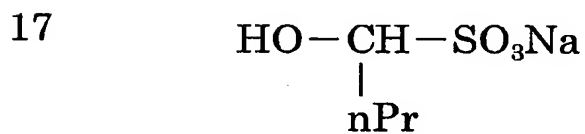
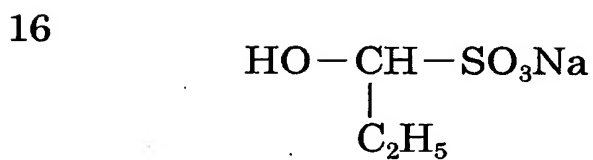
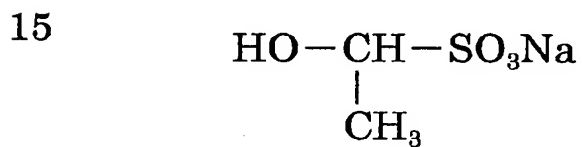
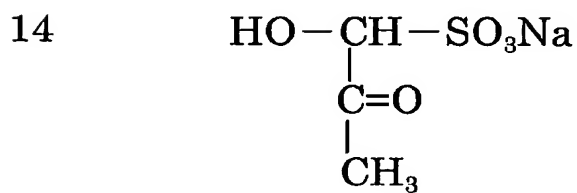
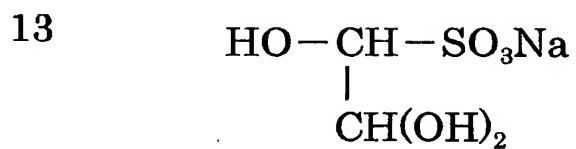
[0018]

[Chemical Formula 4]



[0019]

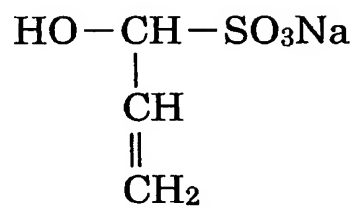
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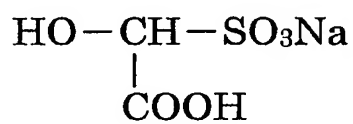
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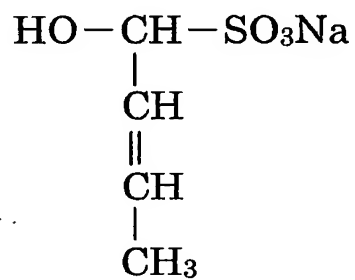
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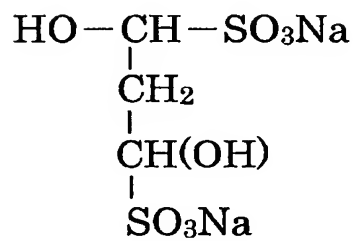
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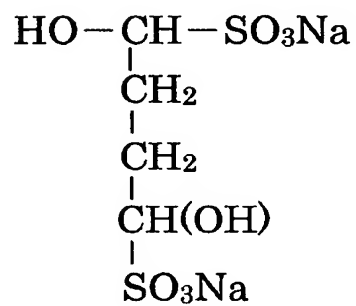
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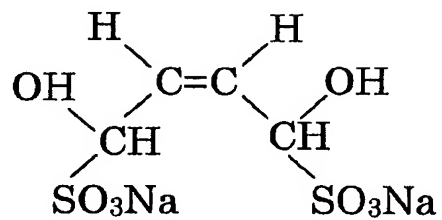
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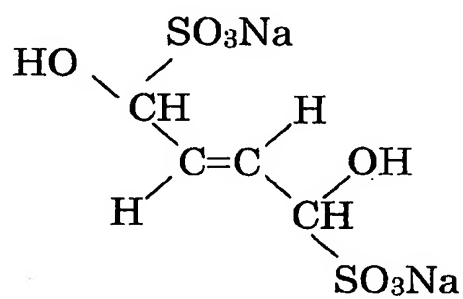
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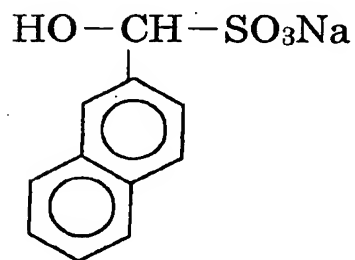
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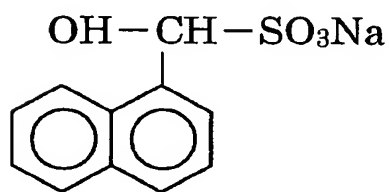
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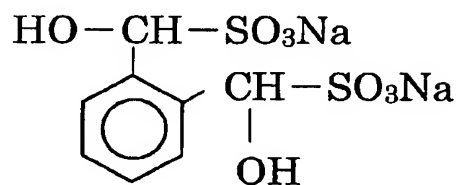
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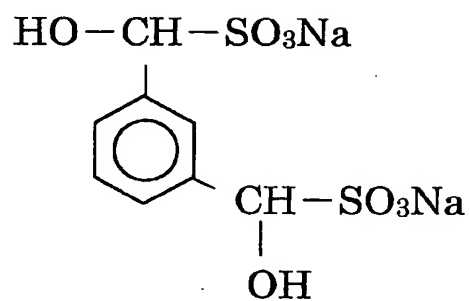
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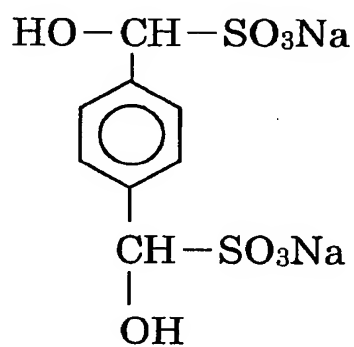
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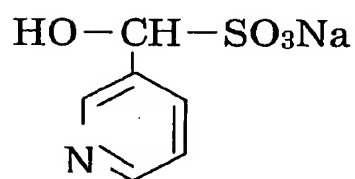
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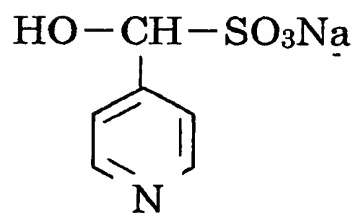
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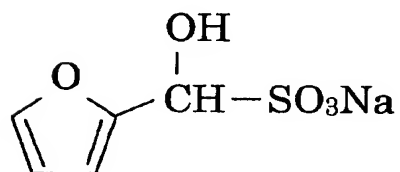
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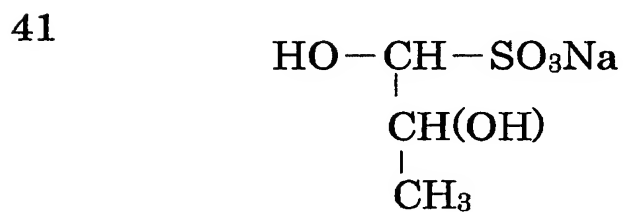
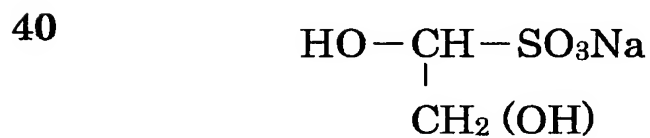
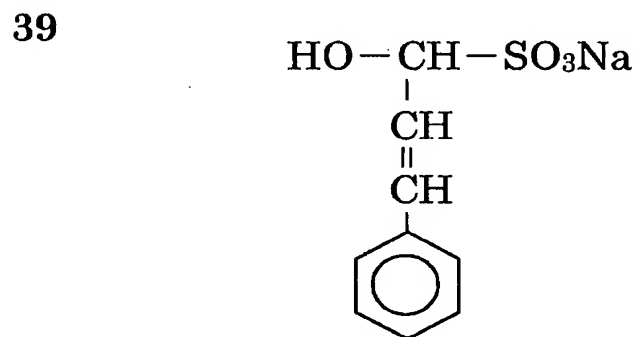
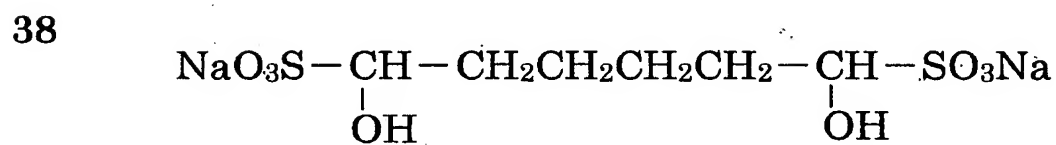
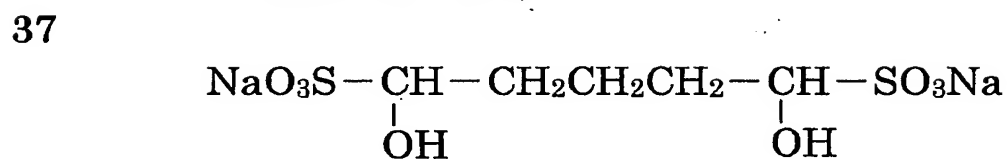
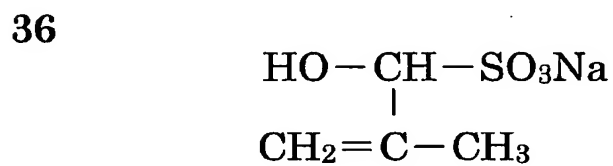
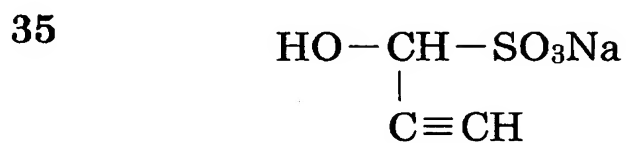
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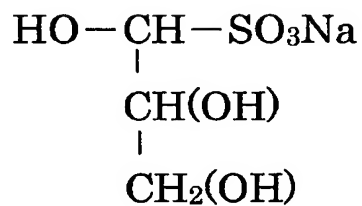


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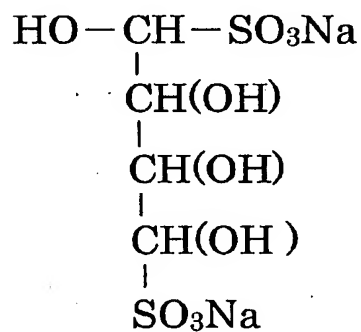


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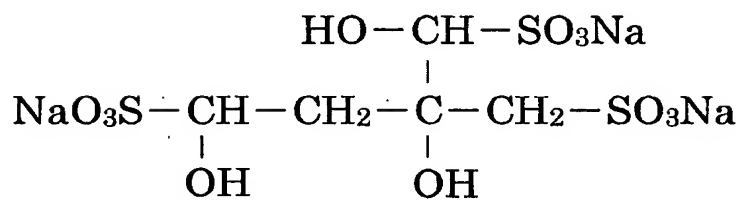
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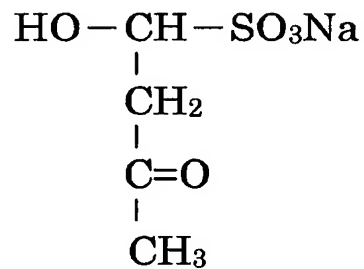
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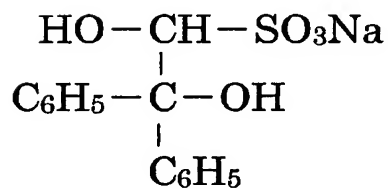
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[0023]

The plating liquid contains preferably 0.001 mol/L to 0.1 mol/L, or more preferably 0.005 mol/L to 0.015 mol/L of the aforementioned hydroxyalkylsulfonic acid or salt thereof. If the content is less than 0.001 mol/L, the gold deposition rate will be slower, while more than 0.1 mol/L is undesirable because of the greater likelihood of bath decomposition and plating irregularities.

[0024]

Examples of the amine compound include compounds having primary and secondary amines in one molecule such as diethylenetriamine, triethylenetetramine, tetraethylenepentamine, pentaethylenehexamine, monoamine compounds such as methylamine, ethylamine, propylamine, butylamine, pentanamine, hexanamine, and diamine compounds such as diaminemethylenediamine, ethylenediamine, propylenediamine, butylenediamine, pentanediamine and hexanediamine.

Other examples include aromatic amines having an aromatic ring such as benzene ring or other bound to the aforementioned compounds, and aromatic amine compounds such as aniline having directly-bound amino groups.

Other examples include aminocarboxylic acids such as glycine and alanine.

Moreover, compounds having a high-polarity substitutional group such as a hydroxyl group, a carboxyl group or a sulfonic acid group bound to the aforementioned compounds or compounds forming a salt such as hydrochlorides can also be used to improve solubility in water.

[0025]

The content of the amine compound in the plating liquid is preferably 0.0005 mol/L to 0.2 mol/L or more preferably 0.003 mol/L to 0.03 mol/L. If the content is less than 0.0005 mol/L, the deposition rate will be slower, while more than

0.2 mol/L is not desirable because the bath becomes unstable.

[0026]

Moreover, sodium dihydrogen phosphate, sodium borate or the like may be added as necessary to the electroless gold plating liquid of the present invention as a pH buffer.

[0027]

In addition, the plating liquid of the present invention can also contain an aminocarboxylic acid compound or a carboxylic acid compound as a complexing agent in addition to the aforementioned mixed complexing agent. Examples of aminocarboxylic acid compounds include ethylenediaminetetraacetic acid (EDTA), hydroxyethylethylenediaminetriacetic acid, dihydroxyethylethylenediamine diacetic acid, propanediamine tetraacetic acid, diethylenetriamine pentaacetic acid, triethylenetetramine hexacetic acid, glycine, glycylglycine, glycylglycylglycine, dihydroxyethylglycine, iminodiacetic acid, hydroxyethyliminodiacetic acid, nitrilotriacetic acid, nitrilotripropionic acid, or alkali metal salts, alkali earth metal salts or ammonium salts thereof and the like. Rochelle salt and the like are also examples of carboxylic acid compounds.

[0028]

It is desirable from the standpoint of gold deposition rate, plating coat appearance and bath stability that the pH of the gold plating liquid of the present invention be 5 to 9, and a pH of 6 to 8 is particularly desirable.

Alkaline compounds such as potassium hydroxide, sodium hydroxide and ammonia, and acid compounds such as sulfuric acid and phosphoric acid can be used to adjust pH.

It is desirable from the standpoint of bath stability and gold deposition rate that the gold plating liquid of the present invention be used at a bath temperature of 60 to 90°C.

[0029]

When plating with the gold plating liquid of the present invention, the material to be plated such as a printed circuit board is dipped in the bath. The material to be plated has preferably first been plated with a nickel undercoat or a subsequent displaced gold coat, and when plating is performed by using the gold plating liquid of the present invention, an adequate deposition rate for practical use is obtained, and the resulting gold plate coat has a good appearance. Plating liquid stability, which was a problem in the past, is also excellent.

[Examples]

[0030]

Preferred embodiments of the present invention are explained with the following examples and comparative examples.

[0031]

As shown in Table 1 below, a 70 μm thick rolled copper foil (glossy on both sides, total area 15.8 cm^2) which is used as the test piece is immersed for 5 minutes at about 45°C in PB-242D acidic degreaser (made by Nikko Metal Plating K.K.) to remove oxides and organic substances such as a rolling oil which might be adhering somewhat to the surface of the copper foil. Next, it is immersed for 1 minute in 50°C hot water to efficiently remove the acidic degreaser from the copper foil, and then water washed for about 1 minute. It is then immersed for 45 seconds at about 25°C in a sodium persulfate solution (sodium persulfate 100 g/L, 96% sulfuric acid 20 mL/L) to bare a fresh copper foil surface, and water washed for 1 minute. Then it is dipped for 2 minutes at room temperature in a sulfuric acid solution (96% sulfuric acid 30 mL/L), and water washed for 1 minute. It is then dipped for 30 seconds at about 25°C in a hydrochloric acid solution (35% hydrochloric acid 100 mL/L) to prevent sulfuric acid contamination.

[0032]

This is then immediately immersed for 2 minutes at about 25°C in KG-522 Pd activator liquid (made by Nikko Metal Plating K.K.), and water washed for 1 minute. It is then immersed for 10 seconds at about 25°C in a sulfuric acid solution (96% sulfuric acid 30 mL/L) to prevent contamination by the Pd activator liquid, and water washed for 1 minutes. Electroless nickel plating (using KG-530 made by Nikko Metal Plating K.K.) is then performed for 20 minutes at 90°C to prevent oxidation of the copper foil surface, depositing nickel to a thickness of about 7 μm on both surfaces of the copper foil. This is then water washed for 1 minute, and dipped for 20 minutes at 80°C in CF-500S electroless displacement plating liquid which doesn't cause pin hole corrosions on the nickel plated surface (made by Nikko Materials K.K.) to deposit gold to a thickness of 0.05 μm on the nickel surface. This is then water washed for 1 minute, and electroless gold plating of the present invention is performed. Drying before and after electroless gold plating is done for the purpose of measuring the thickness of the gold plate by the weighing method, and is not necessary for actual processing.

[0033]

The measuring methods and evaluation methods adopted in the electroless plating methods of the following examples and comparative examples are as follows.

[Method to measure plating rate]

The weights of the test piece before and after the electroless gold plating test were measured with an electronic force balance capable of measuring to 0.1 mg, and the thickness of the gold was calculated. The calculation formula is as follows:

Gold thickness (μm) = (final weight (g) - initial weight (g)) \times 10,000 ($\mu\text{m}/\text{cm}$) \div density of gold (19.3 g/cm³)

÷ area (15.8 cm²).

[0034]

[Method of evaluating appearance]

Color, irregularities and luster were evaluated visually after the plating test. The color of the test piece after plating was judged to be gold if it had a gold color with the deep metal luster of gold based on the color after electroless displacement gold plating. It was not considered gold if it exhibited any brown or reddish color, but was instead described as yellow-brown or reddish-yellow for example. Gold was the only color seen in the current test results.

The presence or absence of irregularities was evaluated at least as strictly as color judgment. The rolled copper foil already had some irregularities, and irregularities due to electroless nickel plating and electroless displacement gold plating were also distinguished when observing. However, there were judged to be no irregularities in all cases in the current test results.

The test pieces after electroless displacement gold plating exhibited luster, and based on this standard it was judged to be no luster in the observation if the luster was inferior to this, but in the present test results there was luster in all cases.

[0035]

[Liquid color]

The color of the plating liquid was evaluated with the naked eye immediately after completion of the electroless gold plating test. The electroless gold plating liquids of the present invention were clear, but after the test one appeared slightly bluish.

[Bath decomposition]

Immediately after finishing the electroless gold plating test the test piece was removed from the container, the

container was wrapped in wrapping film and stored for a day and night at the test temperature (70 to 85°C), and after removal of the film the occurrence of abnormal gold deposition in the container was examined. As a result, abnormal gold deposition starting from a flaw in the container was seen only in Comparative Example 1. Under the other test conditions, no abnormal gold deposition was observed despite severe storage conditions.

[0036]

[Preparation and testing of electroless gold plating liquids]

Comparative Examples 1 & 2

Plating liquids whose compositions are shown in Table 2 were prepared.

The characteristic differences of Comparative Example 1 were that the plating liquid contained ethylenediamine as a reaction promoter and 8 g/l of hydroquinone as a reducing agent. The plating rate was fast, 0.96 $\mu\text{m}/\text{H}$, but the bath decomposed.

In Comparative Example 2, the bath composition comprised ethylenediamine as a reaction promoter but no reducing agent. Although there was no bath decomposition, the plating rate was 0.45 $\mu\text{m}/\text{H}$, the slowest speed in Tables 2 and 3.

[0037]

Examples 1 through 3

In Example 1 in Table 2, the plating liquid comprised sodium hydroxymethanesulfonate as a reducing agent added to the composition of Comparative Example 2. The plating rate was 0.63 $\mu\text{m}/\text{H}$, 1.4 times than of Comparative Example 2. The bath did not decompose and was extremely stable.

In Example 2 in Table 3, the composition was the same as Example 1 but the bath temperature was 85°C. The plating rate was 0.82 $\mu\text{m}/\text{H}$, 1.82 times that of Comparative Example 2. Even with the bath temperature raised to 85°C the bath did not decompose and was extremely stable.

[0038]

In the bath composition of Example 3, the reducing agent was the same sodium hydroxymethanesulfonate as in Example 1, but glycine was substituted for ethylenediamine as the reaction promoter. The plating rate was 0.64 $\mu\text{m}/\text{H}$, approximately the same as in Example 1, and no bath decomposition occurred. Glycine can be used equally as ethylenediamine as a reaction promoter.

[Table 1]

Patent Examples

Test procedure	Test piece: Cu foil		
	•Chemicals	•Con- ditions	•Processing time
(1) Acid degreasing	PB-242D	40-50°C	5 min
Hot water wash		50°C	1 min
Water wash			
(2) Soft etching	Sodium persulfate	25°C	45 sec
	100g/L		
	Sulfuric acid		
	20 mL/L		
Water wash			
Acid dip	Sulfuric acid		2 min
	30 mL/L		
Water wash			
HCl dip	35% HCl	25°C	30 sec
	100 mL/L		
(3) Activator	KG-522	25°C	2 min
Water wash			
Acid dip	Sulfuric acid	25°C	10 sec
	30 mL/L		
Water wash			
(4) Electroless Ni-P plating	KG-530	90°C	20 min
Water wash		pH 4.5	
(5) Electroless gold displacement plating	CF-500S	80°C	20 min
Water wash		pH 7.0	
Drying			
Initial weighing			
(6) Electroless gold reduction plating	Test	70-85°C	60 min
Drying		pH 7.0	
Final weighing			

[0039]

[Table 2]

Test	Composition	Molecular weight	Comparative Example 1		Comparative Example 2		Example 1	
			g/l	mol/l	g/l	mol/l	g/l	mol/l
Sodium gold sulfite	$\text{Na}_3\text{Au}(\text{SO}_3)_2$	426.06	4 as Au	0.02	4 as Au	0.02	4 as Au	0.02
EDTA-2Na	$\text{C}_{10}\text{H}_{14}\text{N}_2\text{O}_8\text{Na}_2 \cdot 2\text{H}_2\text{O}$	372.24	80	0.215	80	0.215	80	0.215
Potassium sodium tartrate	$\text{C}_4\text{H}_4\text{KNaO}_6 \cdot 4\text{H}_2\text{O}$	282.22	26	0.092	26	0.092	26	0.092
Ethylenediamine	$\text{NH}_2\text{CH}_2\text{CH}_2\text{NH}_2$	60.1	2.5	0.042	2.5	0.042	2.5	0.042
Glycine	$\text{NH}_2\text{CH}_2\text{COOH}$	75.1	0	0.000	0	0.000	0	0.000
HMSNa	$\text{HOCH}_2\text{SO}_3\text{Na}$	138.12	0	0.000	0	0.000	1.4	0.010
Sodium sulfite	Na_2SO_3	126.04	45	0.357	45	0.357	45	0.357
Sodium thiosulfate	$\text{Na}_2\text{S}_2\text{O}_3$	158.11	5.5	0.035	5.5	0.035	5.5	0.035
Hydroquinone	$\text{C}_6\text{H}_4(\text{OH})_2$	110.11	8	0.073	0	0.000	0	0.000
pH			7		7		7	
Bath temp.	°C		70		80		80	
Plating rate								
Appearance	$\mu\text{m}/\text{H}$		0.960		0.450		0.630	
Color			Gold		Gold		Gold	
Irregularities			None		None		None	
Luster			Yes		Yes		Yes	
Liquid color			Clear		Clear		Clear	
Bath decomposition			Yes		None		None	

[0040]

[Table 3]

Test	Composition	Molecular weight	Example 2		Example 3	
			g/l	mol/l	g/l	mol/l
Sodium gold sulfite	$\text{Na}_3\text{Au}(\text{SO}_3)_2$	426.06	4 as Au	0.02	4 as Au	0.02
EDTA-2Na	$\text{C}_{10}\text{H}_{14}\text{N}_2\text{O}_8\text{Na}_2 \cdot 2\text{H}_2\text{O}$	372.24	80	0.215	80	0.215
Potassium sodium tartrate	$\text{C}_4\text{H}_4\text{KNaO}_6 \cdot 4\text{H}_2\text{O}$	282.22	26	0.092	26	0.092
Ethylenediamine	$\text{NH}_2\text{CH}_2\text{CH}_2\text{NH}_2$	60.1	2.5	0.042	0	0.000
Glycine	$\text{NH}_2\text{CH}_2\text{COOH}$	75.1	0	0.000	7.5	0.100
HMSNa	$\text{HOCH}_2\text{SO}_3\text{Na}$	138.12	1.4	0.010	1.4	0.010
Sodium sulfite	Na_2SO_3	126.04	45	0.357	45	0.357
Sodium thiosulfate	$\text{Na}_2\text{S}_2\text{O}_3$	158.11	5.5	0.035	5.5	0.035
Hydroquinone	$\text{C}_6\text{H}_4(\text{OH})_2$	110.11	0	0.000	0	0.000
pH			7		7	
Bath temperature	$^{\circ}\text{C}$		85		80	
Plating rate			0.820		0.640	
Appearance		$\mu\text{m}/\text{H}$				
Color			Gold		Gold	
Irregularities			None		None	
Luster			Yes		Yes	
Liquid color			Clear (faint blue)		Clear	
Bath decomposition			None		None	

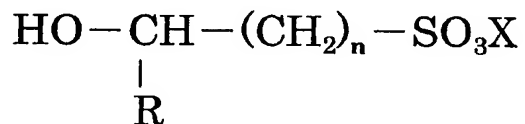
[Designation of the Document] Abstract

[Abstract]

[Problem] To provide an electroless plating liquid which has an adequate deposition speed for practical use without containing any thallium or other heavy metal ions, and which has excellent stability of the plating liquid.

[Means for Solving the Problem] An electroless gold plating liquid comprising a non-cyanide gold salt as a gold salt, an alkali metal salt or an ammonium salt of sulfurous acid and thiosulfuric acid as a metal complexing agent, a hydroxyalkylsulfonic acid or a salt thereof represented by the following general formula as a reducing agent, and an amine compound.

[Chemical Formula 1]



(wherein R represents hydrogen, a carboxyl group, or any of a phenyl group, a naphthyl group, a saturated or unsaturated alkyl group, an acetyl group, an acetonyl group, a pyridyl group and a furyl group which may have a substitutional group, X represents any of hydrogen, Na, K, and NH_4 , and n is an integer between 0 and 4.)

[Selected Figure] None

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Additional Information**

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Address: 10-1, Toranomom 2-chome, Minato-ku, Tokyo

Name: NIKKO MATERIALS CO., LTD.